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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$ $3 O_2 \rightarrow 2 CO_2(g) + 3 H_2O$ $-1411 \text{ kJ mol}^{-1}$ -286 kJ mol^{-1} $-1560 \text{ kJ mol}^{-1}$ (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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8 20 2 2 2 2 o correct reactants and products; [2] correct

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coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO / (unburnt) hydrocarbons; Do not accept methane/CH₄, ethane/C₂H₆, propane/C₃H₈ or butane/C₄H₁₀. Global pollutant: nitrogen oxide/NO / carbon ...

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mol ÷ Ö 1 -1 CH (l) 1 2O(g) 8CO(g) 9HO(l) 8 18 + → + 2 2 2 2 H

=-5470 kJ mol⁻¹ (i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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possible liver damage; [2 max] (b) mild analgesics function by intercepting the pain stimulus at the source / interfere with the production of substances that cause pain/prostaglandins; strong analgesics work by bonding to receptor sites in the brain / prevent the transmission of pain impulses without depressing the central nervous system; [2]

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